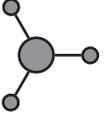
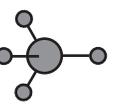
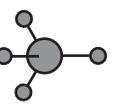
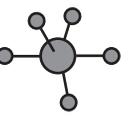
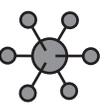


## Molecular Shapes As Developed By The VSEPR Model

<u>Total Pairs</u>	<u>Bonded Pairs</u>	<u>Lone Pairs</u>	<u>Shape</u>	<u>Bond Angle</u>	<u>Structure</u>	<u>Example</u>
1	1	0	Linear	180°		sodium iodide (NaI)
2	2	0	Linear	180°		beryllium fluoride (BeF <sub>2</sub> )
3	3	0	Trigonal Planar	120°		boron fluoride (BF <sub>3</sub> )
3	2	1				
3	1	2				
4	4	0	Tetrahedral	109.5°		methane (CH <sub>4</sub> )
4	3	1	Trigonal Pyramidal	107°		ammonia (NH <sub>3</sub> )
4	2	2	V-Shaped	104.5°		water (H <sub>2</sub> O)
4	1	3	Linear	180°		sodium iodide (NaI)
5	5	0	Trigonal Bipyramidal	120°, 90°		phosphorus pentachloride (PCl <sub>5</sub> )
5	4	1	See-Saw	120°, 90°		sulfur tetrafluoride (SF <sub>4</sub> )
5	3	2	T-Shaped	90°		bromine trifluoride (BrF <sub>3</sub> )
5	2	3	Linear	180°		xenon difluoride (XeF <sub>2</sub> )
5	1	4	Linear	180°		
6	6	0	Octahedral	90°		sulfur hexafluoride (SF <sub>6</sub> )
6	5	1	Square Pyramidal	90°		iodine pentafluoride (IF <sub>5</sub> )
6	4	2	Square Planar	90°		xenon tetrafluoride (XeF <sub>4</sub> )
6	3	3	T-Shaped	90°		
6	2	4	Linear	180°		
6	1	5	Linear	180°		